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NaOH(aq) = 5.0 mL; [NaOH(aq)] = 0.500 mol/L; Required: [H+(aq)], pH Analysis: Before titration, amount of H+ is: H(aq) H(aq) H(aq) 0.350 mmol mL ++ += x = n c VSection 8.7: Acid-Base Titration Tutorial 1 Practice, page 547Complex chemistry concepts are presented in a clear, understandable fashion and key concepts, such as thermodynamics, are treated in greater depth than specified in the curriculum. Nelson Chemistry 12 provides a rigorous, comprehensive, and accurate treatment of all concepts and processes presented in Ontario's chemistry, Grade 12, university Preparation course (SCH4U).Buy Nelson Chemistry 12 Solutions ManualChemistry 12. Here you will find course resources for SCH 4U Chemistry course at UFA ... Nelson has provided access to the online

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reaction is at equilibrium. (b) Given: $[HI(g)] = 0.20 \text{ mol/L}$; $[H_2(g)] = 0.15 \text{ mol/L}$; $[I_2(g)] = 0.090 \text{ mol/L}$; $K = 0.020$ Required: Q ; direction of reaction to reach equilibrium Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.15)(0.090)}{(0.20)^2}$ $Q = 0.34$ Section 7.5: Quantitative Changes in Equilibrium Systems ... Download Nelson Chemistry 12 Solutions Manual - test.enableps.com book pdf free download link or read online here in PDF. Read online Nelson Chemistry 12 Solutions Manual - test.enableps.com book pdf free download link book now. All books are in clear copy here, and all files are secure so don't worry about it. Nelson Chemistry 12 Solutions Manual - Test.enableps.com ... Solutions Nelson Chemistry 12 Chapter 7 Solutions This is likewise one of the

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Statement: The value of Q is 0.020. Q is equal to K, so the reaction is at equilibrium. (b) Given: $[HI(g)] = 0.20 \text{ mol/L}$; $[H_2(g)] = 0.15 \text{ mol/L}$; $[I_2(g)] =$

0.090 mol/L; $K = 0.020$ Required: Q ;
 direction of reaction to reach equilibrium
 Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} =$
 $(0.15)(0.090) / (0.20)^2$ $Q = 0.34$

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