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Thermochemistry 5-5
Solution: 4.18 J (1.00 g)
 $q = mc\Delta T = 1.00 \text{ g} \times 4.18 \text{ J/g} \cdot \text{C} \times 10.5 \text{ C}$
 $q = 43.9 \text{ J}$

Statement: The

amount of energy absorbed by 1.00 g of water, H₂O(l), when its temperature is raised 10.5 °C is 43.9 J. 36.

Given: V HCl(aq) = 100.0 mL; V

NaOH(aq) = 100.0 mL;

T initial = 23.5 °C;

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Indicate whether the sentence or statement is true or false. 1.

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exothermic reactions, the reactants have more kinetic energy than the products.

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om Chapter 1, Organic

Compounds, introduces

the historical concept

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as those compounds

produced by living

things. Students then

learn that the current

concept of organic

chemistry is the study

of carbon compounds.

Each section of the

chapter addresses a

different group of

organic compounds, related by structure. TEACHER'S RESOURCE - NelsonSection 6.5: Rate Law. Tutorial 1 Practice, page 380 1. (a) The order of reaction with respect to A is 2. When the initial concentration of A is doubled, the rate increases by a factor of 4, or 2², so, the exponent of the concentration of reactant A is 2. The order of reaction with respect to B is 1. When the initial concentration of B is doubled, the rate is doubled, or multiplied by 2¹, so the exponent of the concentration of reactant B is 1. Section 6.5: Rate Law - Pre University Courses Nelson Education > School > Secondary Science > Chemistry 12 >

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 Statement: The amount of energy absorbed by 1.00 g of water, $\text{H}_2\text{O}(\text{l})$, when its temperature is raised 10.5°C is 43.9 J. 36.

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