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solution after 3 hours? (Delhi & All India 2009) Answer: Given :  $[R]_0 = 0.10 \text{ M}$ ,  $t = 3 \text{ hrs} = 180 \text{ min}$

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d. Write the rate law for the overall reaction.  $\text{rate} = k [A]^2 [B]^2$

9. Consider the following mechanism.  $O_3 \rightarrow O_2 + O$  (fast)  $O_3 + O \rightarrow 2 O_2$  (slow)

a. Write the overall balanced chemical equation.  $2 O_3 \rightarrow 3 O_2$

b. Identify any intermediates within the mechanism.  $O$

c. What is the order with respect to each reactant?

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 $\text{O}_3 \rightarrow \text{O}_2 + \text{O}$  (fast)  
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 a. Write the overall balanced chemical equation.  
 $2 \text{O}_3 \rightarrow 3 \text{O}_2$   
 b. Identify any intermediates within the mechanism.  
 c. What is the order with respect to each

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 1. Given the half-life for either a first or second order reaction, calculate the time-dependence of the concentration of a reactant for all times.  
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Question 18. Show that for a first order reaction the time required for 99% completion of a reaction is twice the time required to complete 90% of the reaction. (C.B.S.E.Outside Delhi 2013)

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taking place. (D) none of the above.  
 Answer.

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1. The rate of a chemical reaction tells us about. the reactants taking part in the reaction; the products formed in the reaction; how slow or fast the reaction is taking place; none of the above; Answer: (c) 2. In the rate equation, when the concentration of reactants is unity then the rate is equal to . specific rate constant; average rate constant  
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