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Section 2 Chemical Formulas and Equations

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1 d. Cu in Cu_2S 6 e. Cr in Na_2CrO_4 5 f. N in HNO_3

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 Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on each Fe ion.

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Chemical Formulas And Compounds Chapter 7 Review Section 1
 Chapter 9 Practice Problems: Chemical Names and Formulas. Section 9.1: Naming Ions. 1. What is the charge on the ion typically formed by each element? a. oxygen c. sodium e. nickel, two electrons lost . b. iodine d. aluminum f. magnesium. 2. How many electrons does the neutral atom gain or lose when each ion forms? a. Cr^{3+} c. Li^+ e. Cl^- b. P_3^- ...

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 Step 1: Write the prefix of the first element (if the first element is single, don't put a

prefix) Step 2: Write the name of the first element Step 3: Write the prefix of the second element Step 4: Write the root of the second element Step 5: Add -ide Ex) PCl_3 Phosphorus trichloride (Tri-Chlor-Ide) Ex) S_4F_6

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 Science, Chemical Reactions, Section 1 Chemical Formulas and Equations. STUDY. PLAY. Matter can change physically or chemically; a process that produce a chemical change is a chemical ____ reaction. A chemical ____ is a shorthand form of what reactants are used and what products are formed in a chemical reaction.

Science, Chemical Reactions, Section 1 Chemical Formulas ...
 chemical formula is a shorthand way to use chemical symbols and numbers to represent a substance. A chemical formula shows how many atoms of each kind of element are present in a molecule. The chemical formula for water, shown in Figure 1, is H_2O . This formula tells you that one water molecule is made of two hydrogen atoms and one oxygen atom.

Section 2 Chemical Formulas and Equations
 Chapter 7 Section 1 Chemical Names and Formulas Significance of a Chemical Formula, continued •The chemical formula for an ionic compound represents one formula unit— the simplest ratio of the compound's positive ions (cations) and its negative ions (anions). -example: aluminum sulfate — $\text{Al}_2(\text{SO}_4)_3$...

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c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions

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chemical formula is a shorthand way to use chemical symbols and numbers to represent a substance. A chemical formula shows how many atoms of each kind of element are present in a molecule. The chemical formula for water, shown in Figure 1, is H_2O . This formula tells you that one water molecule is made of two hydrogen atoms and one oxygen atom.

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Section 9.1: Naming Ions. 1. What is the charge on the ion

typically formed by each element? a. oxygen c. sodium e. nickel,

two electrons lost . b. iodine d. aluminum f. magnesium. 2. How

many electrons does the neutral atom gain or lose when each ion

forms? a. Cr^{3+} c. Li^+ e. Cl^- b. P^{3-} ...

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