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Reaction. Process in which
one or more substances
are converted into new
substances. Reactant.
Substance that enters a
chemical reaction.
Product. Substance
formed during a chemical

reaction. Balanced
chemical
equation.Chemistry
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group number.Chapter 9 - Study Guide Chemistry Flashcards | QuizletCHEMISTRY- Chapter 9 Study Guide Define acid and base according to Arrhenius and Bronsted-Lowry. Arrhenius definition, an acid is a substance that produces hydrogen ions (H^+) in aqueous solution. A base is a substance that produces hydroxide ions (OH^-) in aqueous solution.Chapter 9 Study Guide - CHEMISTRY Chapter 9 Study Guide ...118 Study Guide for An Introduction to Chemistry

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type of orbital is (11) 52
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 definition, an acid is a
 substance that produces
 hydrogen ions (H +) in
 aqueous solution. A base
 is a substance that
 produces hydroxide ions

(OH-) in aqueous solution.

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monatomic ions. consist
of a single atom with a
positive or negative
charge. how is the ionic
charge of Group 1A, 2A,
or 3A determined. they
lose electrons-form
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charges equal to their
group number.

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structures give atomic
connectivity: they tell us
which atoms are
physically connected to
which atoms. •The shape
of a molecule is
determined by its bond
angles.

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textbook, read about
aqueous solutions,
reactions that form
precipitates, reactions
that form water, and
reactions that form gases.
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choice that best
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chemical equation.

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orbitals that result are
called (10) orbitals. An

example of a molecule
that has this type of
orbital is (11) 52
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