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continued ____21. Metal
X replaces the ions of

metal Y from solution, but
it cannot replace the ions
of metal Z from solution.
The order these metals
should have in the activity
series (from top to
bottom) is a. x, y, z b. z, x,
yAssessment Chapter Test
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continued 17. A substance
combines with oxygen,
releasing a large amount
of energy as heat and
light, in a(n) . 18. The
decomposition of a
substance by an electric
current is called . 19. A(n)
orders the elements by
the ease with which they
undergo certain chemical
reactions. 20. Assessment
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Chapter Test B, continued
27. Distinguish between
ionic crystals and metallic
crystals. PART V Write the
answers to the following
questions on the line to
the left, and show your
work in the space
provided. The molar
enthalpy of fusion for
water is 6.008 kJ/mol.
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Test Chapter: States of
Matter In the space
provided, write the letter
of the term or phrase that

best completes each
statement or best
answers each question.
____ 1. The kinetic-
molecular theory of gases
assumes that the particles
of an ideal gas are
separated by great
distances. This implies
that the gas Assessment
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B, continued 16. The
measure of the ability of
an atom in a chemical
compound to attract
electrons from another
atom in the compound is
called _____. 17. The
energy required to
remove one electron from
an atom is called its _____.
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_____ 7. The melting
points of ionic compounds
are higher than the
melting points of
molecular compounds
because a. ionic
substances tend to
vaporize at room

temperature. b. ionic substances are brittle. c. attractive forces between ions are greater than the attractive ...Assessment Chapter Test A Modern Chemistry 111 Chapter Test Name Class Date Chapter Test B, continued 18. Dissolution processes with negative enthalpies of solution are processes. 19. is the solution process with water as the solvent. 20. Liquid solutes and solvents that are not soluble in each other are. 21. A solute molecule that is surrounded by solvent molecules is

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Building Blocks of Matter - Ms. Sheehan's ...MODERN CHEMISTRY The atomic number of nickel-60 is 28. How many neutrons does this iso- tope have? Carbon-14 has 8 neutrons. What is the atomic number of carbon-14? 11 CHAPTER 3 TEST HR W material copyrighted under notice appearing earlier in this work. www.srvhs.org Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the

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Chapter Test B, continued
 16. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____.
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Chapter Test Name Class

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 21. A solute molecule that
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