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ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C_3H_4 ? 2 mol O₂:1 mol H₂O c. What is the mole ratio of O₂ to H₂O
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in air to form Mg₃N₂. The pure oxygen should be in excess. 5. a. 50 mol HI 6. a. 15.8 Holt Chemistry Chapter 9: Stoichiometry - Practice Test ... Chapter 9 Stoichiometry Multiple Choice Answers Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas. Chapter 9: Standard Review Worksheet Chapter 9 Review Stoichiometry Answers CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1.

Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$
 a. What is the value of the coefficient x in this equation?
 b. What is the molar mass of C_3H_4 ?

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equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO$

$2(g) + 2H_2O(g)$ 4 a. What is the value of

the coefficient x in this equation? 40.07

g/mol b. What is the molar mass of C_3H_4 ?

2 mol O_2 :1 mol H_2O c. What is the

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Identify a chemical change. Relate the

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