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Stoichiometric
Calculations ...CHAPTER 9
REVIEW Stoichiometry
SECTION 3 PROBLEMS
Write the answer on the
line to the left. Show all
your work in the space
provided. 1. 88% The
actual yield of a reaction
is 22 g and the theoretical
yield is 25 g. Calculate the
percentage yield. 2. 6.0
mol of N₂ are mixed with

12.0 mol of H₂
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Stoichiometry Lesson
Starter $\text{Mg(s)} + 2\text{HCl(aq)} \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2(\text{g})$ • If
2 mol of HCl react, how
many moles of H₂ are
obtained? 1 mol H₂ •
How many moles of Mg
will react with 2 mol of
HCl?Chapter 9
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 Stoichiometry Objective •
 Define stoichiometry. •
 Describe the importance
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 mole ratio relating two
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 equation. Chapter 9
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 SECTION 9-3 PROBLEMS
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actual yield of a reaction
 is 22 g and the theoretical
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 of N_2 are mixed with 12.0
 mol of H_2 according to
 the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$
 N_2 ; 2.0 mol a.4798 CHAP 9
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 Composition
 stoichiometry deals with
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 elements in compounds. •
 Reaction stoichiometry

involves the mass
 relationships between
 reactants and products in
 a chemical
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Stoichiometry CHAPTER 9
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SECTION 9. 74 SECTION
9-1 REVIEW MODERN
CHEMISTRY CHAPTER 9
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SECTION 2 PROBLEMS
Write the answer on the
line to the left. Show all.
Reviewing Concepts
CHAPTER 11 REVIEW Key
Equations 11.1 11.2 U g
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9.1: Team Learning
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individual coefficient does
not tell us anything. What
is important is the ratio
between the reactants
and products. For
example, suppose we
were going to make
cookies and a recipe told
us to use two eggs, some
butter, some flour (etc.)
and we would make some
cookies. The fact
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Quizlet Chapter 9 focuses
on reaction stoichiometry:
using a balanced chemical
equation to calculate the
number of grams, moles,
or particles of
reactants/products
involved in a chemical
reaction. Students had an
introduction to
composition stoichiometry
in Chapter 3 and will now
move on to some more
difficult problems. Chapter
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 Stoichiometry

Stoichiometry comes from
 the Greek ... SECTION
 REVIEW 1. What is
 stoichiometry? 2. How is a
 mole ratio from a reaction
 used in stoichiometric
 problems? 3. For each of
 the following chemical
 equations, write all
 possible mole ratios. a.
 $2\text{HgO}(s) \rightarrow 2\text{Hg}(l)$
 $+ \text{O}_2(g)$ CHAPTER 9
 StoichiometryModern
 Chemistry 73
 Stoichiometry CHAPTER 9
 REVIEW Stoichiometry
 SECTION 1 SHORT
 ANSWER Answer the
 following questions in the
 space provided. 1. _____

The coefficients in a
 chemical equation
 represent the (a) masses
 in grams of all reactants
 and products. (b) relative
 number of moles of
 reactants and products.
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... Stoichiometry. 2.
Limiting reagents and
percent yield. NOTES:
Stoichiometry is the
calculation of chemical
quantities from balanced
equations. ... 2 (excess
reagent) will remain by
determining how many
moles of hydrogen will be
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to use mole ratios, molar
masses, conversions,
limiting reactants, and
percent yield to ...
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the mass relationships of
elements in compounds. •
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relationships between
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a chemical reaction.
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Introduction to
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Starter $\text{Mg(s)} + 2\text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$ • If
2 mol of HCl react, how
many moles of H_2 are
obtained? 1 mol H_2 •
How many moles of Mg
will react with 2 mol of
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left. Show all. Reviewing
Concepts CHAPTER 11
REVIEW Key Equations
11.1 11.2 U g mgh E K U g
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CHAPTER 9 REVIEW
Stoichiometry SECTION 3
PROBLEMS Write the
answer on the line to the
left. Show all your work in
the space provided. 1.
88% The actual yield of a
reaction is 22 g and the
theoretical yield is 25 g.
Calculate the percentage
yield. 2. 6.0 mol of N_2 are
mixed with 12.0 mol of H

Section 1 Introduction to Chapter 9

Stoichiometry

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Chapter 9 Section 1

Introduction to

Stoichiometry Objective •

Define stoichiometry. •

Describe the importance of the mole ratio in

stoichiometric

calculations. • Write a

mole ratio relating two substances in a chemical equation.

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stoichiometry? 2. How is a

mole ratio from a reaction

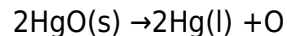
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problems? 3. For each of

the following chemical

equations, write all

possible mole ratios. a.



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Stoichiometric

Calculations ...

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Chapter 9 focuses on

reaction stoichiometry:

using a balanced chemical

equation to calculate the

number of grams, moles,

or particles of

reactants/products

involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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Stoichiometry SECTION

9-3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent

yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g) N₂; 2.0 mol a.

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Stoichiometry Section
2**

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represent the (a) masses and products. (b) relative number of moles of
in grams of all reactants reactants and products.

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