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vessel and heated to 500. K. The following reaction takes place.  $2 \text{IBr (g)} \rightleftharpoons \text{I}_2 \text{(g)} + \text{Br}_2 \text{(g)}$  At equilibrium, the partial pressure of  $\text{Br}_2$  is 3.01 atm. Calculate  $K_p$  for this reaction. Calculate  $K_c$  for the same reaction. Q11. Consider the reaction  $2 \text{SO}_3 \text{(g)} \rightleftharpoons 2 \text{SO}_2 \text{(g)} + \text{O}_2 \text{(g)}$  ...CHM 112 Chapter 13 Worksheet - nvcc.edu CHM 112 Chapter 13 Worksheet: Chemical Kinetics Name: \_\_\_\_\_ Q1. The reaction  $2 \text{NO}_2 \text{(g)} \rightarrow 2 \text{NO (g)} + \text{O}_2 \text{(g)}$  was studied at 383 °C and the following data was obtained. Time (s) [NO<sub>2</sub>](M) 0.0 0.100 5.0 0.017 10.0 0.0090 15.0 0.0062 20.0 0.0047 Plot this data and determine if the reaction is first or second order. CHM 112 Worksheet Chapter 12 - nvcc.edu CHM 112 Chapter 15 Terms study guide by redskinjoe includes 28 questions covering vocabulary, terms and more. Quizlet flashcards, activities and games help you improve your grades. CHM 112 Chapter 15 Terms Flashcards | Quizlet Chapter 15 Worksheet 1. Bronsted-Lowry definition of acids and bases, pH scale, strong acids and bases. Arrhenius' definition of acids and bases: Acid: Substance that produces aqueous H<sup>+</sup> (protons) when dissolved in water. Hydrochloric acid:  $\text{HCl(aq)} \rightarrow \text{H}^+\text{(aq)} + \text{Cl}^-\text{(aq)}$   $K_a$  is large; strong acid. Chapter

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temperature and pressure are 273 K (0 C, 32 F) and a pressure of 1 atm. 2. What is the volume of 0 ...CHM 113 Ch 10 wksht - CHM 113 Hendrickson Kira Wagner ...Learn chm chapter 14 with free interactive flashcards. Choose from 500 different sets of chm chapter 14 flashcards on Quizlet. CHM 112 Chapter 15 Worksheet: Acids and Bases Name: The ... Properties Of Acids And Bases - Displaying top 8 worksheets found for this concept.. Some of the worksheets for this concept are Acids bases work, Properties of acids and bases, 11 0405 acids bases salts wkst, Acidsbases ph work, Properties of acids and

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$2 / T_2 P_1 V_1 / T_1 = P_2 V_2 / T_2$   $PV = nRT$   
1 atm = 1.01 x 10<sup>5</sup> Pa Density = mass/volume  
1. What are the conditions for STP? Standard temperature and pressure are 273 K (0 C, 32 F) and a pressure of 1 atm.  
2. What is the volume of 0 ...

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Chapter 15 Worksheet 4 (ws15.4) Acid-Base Properties of Salt Solution, Polyprotic Acids. All acid-base reactions produce another acid and another base. The reaction between the stronger acid and the stronger base is called a "neutralization reaction". Neutralization reactions produce a salt (an ionic compound) and (sometimes) water.

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