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 $A(g) = 2B(g) + C(g)$
 $A(g) = 2X(g) + C(g)$
 $K_c = 3.35$
 $B(g) = X(g)$
 $K_c = 22.0$
 Express your answer using three significant figures.
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Equilibrium
 Chemical Equilibrium Section 18.1 Equilibrium: A State of Dynamic Balance
 In your textbook, read about chemical equilibrium. Complete each statement.
 1. When a reaction results in almost complete conversion of reactants to products, chemists
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 Question: 4-1 Experiment 4 CHEMICAL EQUILIBRIUM DISCUSSION (Reference: Equipment Guide: CCD Array Spectrometer. Measurement, Uncertainty. And Significant Figures. Elementary Statistical Concepts. BLBMWS, Ch 15 Sections 1-3,5-6.)
 In This Experiment, You Will Study The Reaction:
 $(1) Fe^{3+}(aq) + HSCN(aq) \rightleftharpoons Fe(SCN)^2+(aq) + H^+(aq)$
 When Solutions Containing
 Festions ...
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 The product of the forward chemical reaction is HI, for the equilibrium expression:
 $K_{eq} = \frac{[HI]^2}{[H_2][I_2]}$
 1
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 In your textbook, read about determining equilibrium constants.
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almost complete conversion of reactants to products, chemists

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10. Possible answers include: white with blue lines, flat, thin, 1-2 g, flammable, solid. 11. Possible answers include: red/green/red & blue, cylindrical solid filled with liquid which has a high concentration of gas, 355 mL, pressure increases if shaken. 12. Possible answers include: color, rectangular. solid, made of paper and cardboard ...

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Chemical equilibrium can be disturbed and shifted away from equilibrium by: __, __, __ Le Chatelier's Principle ____ states: If an external stress is applied to a system at equilibrium, the system adjusts in such a way that the stress is partially offset as the system reaches a new equilibrium position.

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Question: 4-1 Experiment 4 CHEMICAL EQUILIBRIUM DISCUSSION (Reference: Equipment Guide: CCD Array Spectrometer. Measurement, Uncertainty. And Significant Figures. Elementary Statistical Concepts. BLBMWS, Ch 15 Sections 1-3,5-6.) In This Experiment, You Will Study The Reaction: $(1) Fe^{3+}(aq) + HSCN(aq) \rightleftharpoons Fe(SCN)_2(aq) + H^+(aq)$ When Solutions Containing Festions ...

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Find the equilibrium concentration if the equilibrium constant is given Find the equilibrium concentration of one of the reactants in a reaction when the equilibrium expression is a perfect square. Find the equilibrium concentration when the equilibrium constant is given and the equilibrium expression is NOT a perfect square; you will use the quadratic equation.

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Online Library Chemistry Study Guide Answers Chemical Equilibrium Free Study Guide Chemistry Stoichiometry Answer Key Study Guide Chemistry Stoichiometry Answer The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. ____ 2.

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The product of the forward chemical reaction is HI, for the equilibrium expression: K_{eq} Chemistry: Matter and Change [HI]² [H₂]² [I₂] 1 Study Guide Name ____ Date ____ Class ____ 13 17 Section 17.1 continued In your textbook, read about determining equilibrium constants.